

EASTERN UNIVERSITY, SRI LANKA THIRD EXAMINATION IN SCIENCE – 2003/2004 (Repeat) SECOND SEMESTER (June/July-2005) CH 305 ORGANOMETALLIC AND NON-AQUEOUS SOLVENTS

Answer all questions

Time: One hour

1) a) Indicate the monohapto, dihapto,trihapto, tetrahapto and pentahapto ligants present in the following organometallic compounds.



(20 marks)

b) pKa value of $\underline{\mathbf{A}}$ is 11.25 where as that of $\underline{\mathbf{B}}$ is 7.35. Explain this observation.



(15 marks)

c) Give the systematic names of the following organometallic compounds.



(15 marks)

$$CH = CH \xrightarrow{H_2Fe(CO)_4} \underline{A} \xrightarrow{HCo(CO)_3} B$$

i) Give the structures of \underline{A} and \underline{B} .

d)

ii) Give the mechanism involved in the conversion of $\underline{\mathbf{A}}$ to $\underline{\mathbf{B}}$.

(20 marks)

(20 marks)

(10 marks)

e) Explain the bonding in transition metal π -allyl complexes.

2) a) Explain why the Co stretching frequency in $[Cr(dien)(CO)_3]$ complex shows a lower value than the Co stretching frequency in $Cr(CO)_6$.

(20 marks)

b) Using ¹H NMR spectroscopy, explain how you would differentiate σ -bonded and π -bonded cyclopentadienyl ligand.

(15 marks)

c) State whether each of the following would act as an acid or a base in liquid HF. BF₃, SbF₅, H₂O, CH₃COOH

(20 marks)

d) Acetic acid acts as a differentiating solvent for strong acids whereas water acts as a leveling solvent for strong acids. Explain this statement.

(20 marks)



d) Give one example for each of the following types of reaction.

i) Solvolysis reaction in liq. NH₃.

ii) Self-ionization of a protic solvent.

iii) Amphoteric reaction in liq. NH₃.

(15 marks)

e) Give reasons for selecting liquid NH3 as a non aqueous solvent.

(10 marks)